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an increase in the concentration of the reactants of reaction $2a + b \rightarrow \text{product}$ doubling the initial concentration of which of the following factors are equal in a reversible reaction when two reactants a and b are mixed to give products c and d in a reaction the threshold energy is equal to the activation energy consider the following reactions in which all the reactants and which of the following statements is correct about the reaction given a reversible chemical reaction having two reactants in a reaction q 6 which information is not conveyed by a balanced chemical equation statement i at equilibrium the concentration of reactants and products are equal

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the number of collisions that takes place per second per unit volume of the reaction mixture is known as collision frequency Z as the concentration of reactants increases the number of reactant molecules

for the reaction $2a + b \rightarrow \text{products}$ it is found that doubling the concentration of both reactants increases the rate by a factor of 8 but doubling the concentration of b alone only doubles the rate

when the concentrations of reactants and products have become constant a reaction is said to have reached a point of equilibrium the consistency of measurable properties such as concentration

when two reactants a + b are mixed to give products c + d the reaction quotient q at the initial of the reaction

the activation energy for the reaction $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ is $209.5 \text{ kJ mol}^{-1}$ at 581 K calculate the fraction of molecules of reactants having energy equal to or greater than activation energy

[click here to get an answer to your question](#) consider the following reactions in which all the reactants and

which of the following statements is correct about the reaction given below $4\text{FeS} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{S}$ the total mass of reactants total mass of the products it follows the law of conservation

a reversible chemical reaction having two reactants in equilibrium if the concentrations of the reactants are doubled then the equilibrium constant will remain the same

[click here to get an answer to your question](#) q6 which information is not conveyed by a balanced chemical equation a physical states of reactants

statement i at equilibrium the concentration of reactants and products remains constant
statement ii at equilibrium the rates of the forward and reverse reactions are equal

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